[illegible] COUNTY PUBLIC SCHOOLS

check one: Teacher Observation ☒ Evaluation ☐

Teacher: Eman Haggag School: Tuscarora High School Date: 12/21/10

Teacher Assignment: Science Education Observer/Evaluator: Dale E. Peters

Number of Observations: 1 Number of Evaluations: _____ Recommended Reappointment: _____

Tenured: ☐ Yes ☒ No If Nontenured: Semester ☐ 1 ☐ 2 ☒ 3 ☐ 4

		S	U	N/A
DOMAIN ONE: PLANNING AND PREPARATION				
1a	Demonstrates knowledge of content and pedagogy	☒	☐	☐
1b	Demonstrates knowledge of students	☒	☐	☐
1c	Selects instructional goals	☒	☐	☐
1d	Demonstrates knowledge of resources	☒	☐	☐
1e	Designs coherent instruction	☒	☐	☐
1f	Assesses student learning	☒	☐	☐
Comments:				
DOMAIN TWO: THE CLASSROOM ENVIRONMENT				
2a	Creates an environment of respect and rapport	☒	☐	☐
2b	Establishes a culture for learning	☒	☐	☐
2c	Manages classroom procedures	☒	☐	☐
2d	Manages student behavior	☒	☐	☐
2e	Organizes physical space	☒	☐	☐
Comments:				
DOMAIN THREE: INSTRUCTION				
3a	Communicates clearly and accurately	☐	☐	☐
3b	Uses questioning and discussion techniques	☒	☐	☐
3c	Engages students in learning	☒	☐	☐
3d	Uses assessment for instruction	☒	☐	☐
3e	Demonstrates flexibility and responsiveness	☒	☐	☐
Comments:				
DOMAIN FOUR: PROFESSIONAL RESPONSIBILITIES				
4a	Reflects on teaching	☐	☐	☐
4b	Maintains accurate records	☐	☐	☐
4c	Communicates with families	☒	☐	☐
4d	Contributes to the school community	☐	☐	☐
4e	Demonstrates professional growth	☒	☐	☐
4f	Shows professionalism	☒	☐	☐
Comments:				

Overall Performance Rating Sat. ☐ Unsat. ☐

The Components of Professional Practice have been reviewed in conjunction with this observation/evaluation: yes

Observer/Evaluator's Signature Dale E Peters Teacher's Signature _____ Date: 1/3/11

Note: 1. T wo or more unsatisfactory marks would indicate an unsatisfactory performance rating
2. The teacher's signature on this observation/evaluation is an indication that the teacher has received the observation/evaluation and has had an opportunity to discuss it with the observer/evaluator; the signature does not necessarily indicate agreement with the content.

Initial and Date (for multiple pages): DEP 1/3/11

ADDENDUM TO OBSERVATION
Chemistry Honors

TEACHER: Eman Haggag
EVALUATOR: Dale E. Peters
SCHOOL: Tuscarora High School
DATE: 12/21/10

DOMAIN ONE: PLANNING AND PREPARATION

Objectives:

Students will be introduced to solutions and the classification of mixtures.

Agenda:

Warm up
Notes
Demonstration
Small-Scale Laboratory
HMK

YOU ARE COMMENDED FOR:

1. Your excellent knowledge of content.

IT IS RECOMMENDED THAT YOU:

1. Continue to develop warm-up questions that provide support for the SAT exam.
2. Provide students with a more descriptive agenda/objective(s) for the day. Write objectives in behavioral format.

DOMAIN TWO: THE CLASSROOM ENVIRONMENT

It is good to see student work displayed in the classroom. Environment is relaxed, good use of humor, yet students are focused, on task, and responsive to Ms. Haggag's questions.

YOU ARE COMMENDED FOR:

1. Having established an obvious classroom routine, and behavioral expectations.
2. Good classroom management skills.

DOMAIN THREE: INSTRUCTION

1. Warm-up: Typically a question from the previous day's content.
2. Homework was collected.
3. Students were shown a video clip from the "Ellen" TV show on the properties of corn starch in solution.
4. Students were provided with guided notes for discussion on solutions.
5. Ms. Haggag used a power point presentation as well as the overhead to present information on solutions, which included

   Characteristics of solutions
   Electrolytes- ionic and covalent (which conducts electricity)
   Ms. Haggag used a demonstration of such for enrichment.
   Solvation- used an atomic model as well as an animation of HOH dissolving a compound for lesson enhancement.
   Factors affecting solution formation

6. Students were asked to work in groups of their choosing to design the protocol for dissolving a cube of sugar in 20 ml of water as fast as possible. Students would receive added time for the use of specific materials provided to them for completing the task. Students had three "test" trials before doing the final trial, which was timed by Ms. Haggag.
7. A chart was set up on the white board for students to place their final trial results (time and what they did to dissolve the sugar cube). As the group results started being posted the groups became more competitive.
8. Throughout the lab portion of the class, Ms. Haggag moved continually around the room observing, clarifying, making suggestions, and keeping an eye on students to ensure they were following safety requirements.
9. Students were required to turn in separate lab write-ups. The fastest time recorded for dissolving the sugar cube was 38 seconds.
10. Homework- worksheet on concentration problems.

YOU ARE COMMENDED FOR:

1. Positive interaction and engagement (question and answer) between you and your students.
2. Your use of humor in presenting material, especially the use of the video clip from the "Ellen" show. A great item for capturing student interest in the day's topic.
3. Good movement around the room to assist students with their questions.

IT IS RECOMMENDED THAT YOU:

1. Reduce the use of the term "OK".
2. Consider the addition of the temperature of warmed water as a factor that will cost students additional seconds of time.

DOMAIN FOUR: PROFESSIONAL RESPONSIBILITIES

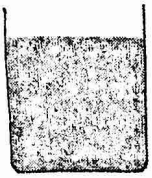
Your willingness to tutor students before, during and after school is appreciated.

YOU ARE COMMENDED FOR:

1. Your professionalism and the respect you display towards your students.
2. Challenging your student's analytical and critical thinking skills.

EVALUATOR'S SIGNATURE Dale Citron DATE 1/3/11

TEACHER'S SIGNATURE _____



SOLUTIONS – Guided Notes

A colloid is _____.

A suspension is _____.

A solution is _____.

The two parts of a solution are _____ and _____, and the solvent is the one in greater quantity.

Solute: _____.

Solvent: _____.

Examples of solutions:

Solute	*Solvent*	*Example*
solid	solid	
solid	liquid	
gas	solid	
liquid	liquid	
gas	liquid	
gas	gas	

Characteristics of solutions:

- Homogeneous mixture
- Soluble – _____.
- Insoluble – _____.
- Immiscible – _____.
- Miscible – _____.

****

An electrolyte is _____.

Most electrolytes are _____, and non electrolytes are _____.

Name _____ Date _____ Class _____

SMALL-SCALE LAB: Electrolytes

Laboratory Recordsheet Use with Section 15.2

SAFETY

Wear safety glasses and follow the standard safety procedures outlined in the Small-Scale Lab Manual.

PURPOSE

To classify compounds as electrolytes by testing their conductivity in aqueous solution.

MATERIALS

- pencil
- paper
- ruler
- reaction surface
- conductivity tester
- chemicals shown in grid below
- water
- micropipet or dropper
- conductivity probe (optional)

PROCEDURE

On a separate sheet of paper, draw a grid similar to the one below. Make each square 2 cm on each side. Place a reaction surface over the grid and place a few grains of each solid in the indicated places. Test each solid for conductivity. Then add 1 drop of water to each solid and test the wet mixture for conductivity. Be sure to clean and dry the conductivity leads between each test. Use the grid as a data table to record your observations.

$NaCl(s)$	$MgSO_4(s)$
$Na_2CO_3(s)$	Sugar ($C_{12}H_{22}O_{11}$)
$NaHCO_3(s)$	Cornstarch $(C_6H_{22}O_6)_n$
$KCl(s)$	$KI(s)$